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The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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VIBRATIONS AND WAVES

Define stoichiometry and describe its importance. Relate stoichiometry to balanced chemical equations. Identify and solve different types of stoichiometry problems. Calculate the amount of product formed in a chemical reaction when reactants are present in nonstoichiometric

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proportions

**CHAPTER 11:
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Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your

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work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: N

mc06se cFMsr i-vi

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calculation, multiply all the terms on the top together ($175 \times 1 \times 2 \times 224$) and divide by the product of the terms on the bottom ($280 \times 1 \times 1$) If you do the calculation accurately, you should find that you have 280 liters of ammonia gas

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Here are a few videos to assist you in study for the test. These videos will walk you through the solving of a few of the problems, but they will not walk you through all of them. Hopefully you can use these to figure out the other ones that are not covered.

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Unit 3 - Chapter 11 - Honors Chemistry

Answer the questions
about the following
chemical reaction.

sodium + iron(III) oxide

→ sodium oxide +

iron $6\text{Na(s)} + \text{Fe}_2\text{O}_3(\text{s}) \rightarrow$

$2\text{Fe(s)} + 3\text{Na}_2\text{O(s)}$

15. What is a

mole ratio? 16. How is

a mole ratio written?

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